

Charges, Names, and Formulas of Common Ions

Positive Ions (Cations)

+1 Charge

<i>Name</i>	<i>Formula</i>
Ammonium	NH_4^+
Copper(I)	Cu^+
Lithium	Li^+
Potassium	K^+
Silver	Ag^+
Sodium	Na^+

+2 Charge

<i>Name</i>	<i>Formula</i>
Barium	Ba^{2+}
Calcium	Ca^{2+}
Copper(II)	Cu^{2+}
Iron(II)	Fe^{2+}
Lead(II)	Pb^{2+}
Magnesium	Mg^{2+}
Nickel	Ni^{2+}
Strontium	Sr^{2+}
Tin(II)	Sn^{2+}
Zinc	Zn^{2+}

+3 Charge

<i>Name</i>	<i>Formula</i>
Aluminum	Al^{3+}
Iron(III)	Fe^{3+}

+4 Charge

<i>Name</i>	<i>Formula</i>
Lead(IV)	Pb^{4+}
Tin(IV)	Sn^{4+}

Negative Ions (Anions)

-1 Charge

<i>Name</i>	<i>Formula</i>
Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$
Bromide	Br^-
Chloride	Cl^-
Dihydrogen Phosphate	H_2PO_4^-
Fluoride	F^-
Hydroxide	OH^-
Hydrogen Carbonate (bicarbonate)	HCO_3^-
Hydrogen Sulfate (bisulfate)	HSO_4^-
Hypochlorite	ClO^-
Iodide	I^-
Nitrate	NO_3^-
Permanganate	MnO_4^-
Thiocyanate	SCN^-

-2 Charge

<i>Name</i>	<i>Formula</i>
Carbonate	CO_3^{2-}
Chromate	CrO_4^{2-}
Dichromate	$\text{Cr}_2\text{O}_7^{2-}$
Hydrogen Phosphate	HPO_4^{2-}
Oxide	O^{2-}
Oxalate	$\text{C}_2\text{O}_4^{2-}$
Sulfate	SO_4^{2-}
Sulfite	SO_3^{2-}

-3 Charge

<i>Name</i>	<i>Formula</i>
Nitride	N^{3-}
Phosphate	PO_4^{3-}

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